

South Pasadena • AP Chemistry

Name Key
Period Day / /**3 • Molecules & Compounds****Mole Calculations - Difficulty Level 1**

$$1 \text{ mole} = 6.02 \times 10^{23} \text{ molecules} = 22.4 \text{ L (@ STP)}$$

1. Calculate the mass of 1.58 moles
- CH_4
- . [molar mass
- $\text{CH}_4 = 16.0 \text{ g/mol}$
-]

G: 1.58 moles CH_4 D: ? g CH_4

$$\frac{1.58 \text{ moles } \text{CH}_4}{1 \text{ mole } \text{CH}_4} \times \frac{16.0 \text{ g } \text{CH}_4}{1 \text{ mole } \text{CH}_4} = 25.28 \approx \boxed{25.3 \text{ g } \text{CH}_4}$$

2. What volume will 7.29 moles of
- CO_2
- gas occupy at STP?

G: 7.29 moles CO_2 D: ? L CO_2

$$\frac{7.29 \text{ moles } \text{CO}_2}{1 \text{ mole } \text{CO}_2} \times \frac{22.4 \text{ L } \text{CO}_2}{1 \text{ mole } \text{CO}_2} = 163.396 \approx \boxed{163 \text{ L } \text{CO}_2}$$

3. How many molecules are there in a 0.00583 mole sample of
- H_2O
- ?

G: 0.00583 moles H_2O D: ? molecules H_2O

$$\frac{0.00583 \text{ moles } \text{H}_2\text{O}}{1 \text{ mole } \text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{H}_2\text{O}} = 3.51 \times 10^{21} \text{ molecules}$$

4. What mass of
- CO_2
- gas occupies a volume of 100. Liters at STP? [molar mass
- $\text{CO}_2 = 44.0 \text{ g/mol}$
-]

G: 100. Liters CO_2 D: ? g CO_2

$$\frac{100. \text{ Liters } \text{CO}_2}{22.4 \text{ L } \text{CO}_2} \times \frac{44.0 \text{ g } \text{CO}_2}{1 \text{ mole } \text{CO}_2} = 196 \text{ g } \text{CO}_2$$

5. How many molecules are in a 35.0 gram sample of
- H_2O
- ? [molar mass
- $\text{H}_2\text{O} = 18.0 \text{ g/mol}$
-]

G: 35.0 g H_2O D: ? molecules H_2O

$$\frac{35.0 \text{ g } \text{H}_2\text{O}}{18.0 \text{ g } \text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{H}_2\text{O}} = 1.17 \times 10^{24} \text{ molecules}$$

6. What volume will
- 5.25×10^{22}
- molecules of
- CH_4
- occupy at STP?

G: 5.25×10^{22} molecules CH_4

D: ? L

$$\frac{5.25 \times 10^{22} \text{ molecules } \text{CH}_4}{6.02 \times 10^{23} \text{ molecules}} \times \frac{22.4 \text{ L } \text{CH}_4}{1 \text{ mole } \text{CH}_4} = 1.95 \text{ L } \text{CH}_4$$

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