

17 INTERPRETING GRAPHICS
Use with Sections 17.2

Thermodynamic quantities are in kilojoules (kJ) unless otherwise indicated. Molar enthalpies of formation are in kJ/mol. All quantities are at 25°C.

Reaction	Enthalpy	Entropy
$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$	44.0 kJ	118.8 J/K
$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(s)$	-6.01 kJ	-22.0 J/K
$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(s)$	-6.01 kJ	-22.0 J/K
$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$	44.0 kJ	118.8 J/K
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$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$	44.0 kJ	118.8 J/K
$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(s)$	-6.01 kJ	-22.0 J/K

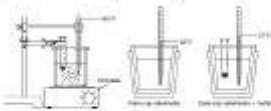


Figure 17.2: Calorimeter Setup

Substance	Specific Heat	Initial Temp
Water	4.184 J/g°C	25.0°C
Substance X	0.50 J/g°C	50.0°C
Substance Y	1.00 J/g°C	25.0°C

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